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Chemistry - Chapter 9 Assessment Answers 9.1 Assessment 3. The positive charge of a group A metal is equal to its group number. The charge of a group A nonmetal is the group number minus 8. 4. from the number of electrons lost 5. -ite or -ate 6. Group 1A metals, 1+; Group 3A (aluminum), 3+ Group 5A nonmetals, 37.

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Solutions Manual Chemistry: Matter and Change • Chapter 9 143  
CHAPTER 9 SOLUTIONS MANUAL Problem-Solving Lab page 294  
Halogen Atomic Radius (ppm) Ionic Energy (kJ/mol) Electronegati-

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vity Fluorine 72 1681 3.98 Chlorine 100 1251 3.16 Bromine 114 1140 2.96 Iodine 133 1008 2.66 Astatine 140 920 2.2 Properties of Halogens 1.

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Answer Section MATCHING 1. ANS: A PTS: 1 DIF: L1 REF: p. 253  
OBJ: 9.1.1 Identify the charges of monatomic ions by using the  
periodic table, and name the ions. STA: Ch.5.a 2.ANS: H PTS: 1  
DIF: L1 REF: p. 254 OBJ: 9 ...

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e. molar mass of  $\text{Fe}(\text{NO}_3)_3 = 241.9 \text{ g}$   $272 \text{ g Fe}(\text{NO}_3)_3$   $1 \text{ mol}$   
 $241.9 \text{ g} = 1.12 \text{ mol Fe}(\text{NO}_3)_3$  f. molar mass of  $\text{FeSO}_4 = 151.9 \text{ g}$   
 $12.7 \text{ mg FeSO}_4$   $1 \text{ g}$   $1000 \text{ mg}$   $1 \text{ mol}$   $151.9 \text{ g} = 8.36 \cdot 10^{-5} \text{ mol}$   
 $\text{FeSO}_4$   $4 \text{ g}$ . molar mass of  $\text{LiOH} = 23.95 \text{ g}$   $6.91 \cdot 10^3 \text{ g LiOH}$   $1 \text{ mol}$   
 $23.95 \text{ g} = 288.5 = 289 \text{ mol LiOH}$  14. a. molar mass of  $\text{CaCO}_3 =$   
 $100.1 \text{ g}$   $2.21 \cdot 10^{-4} \text{ mol CaCO}_3$   $3$   $100.1 \text{ g}$   $1 \text{ mol} = 0 \dots$

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Applying Scientific Methods Date Class CHAPTER ASSESSMENT A  
chemist isolated four samples, A, B, C, and D. She obtained the  
following atomic emission spectra of the samples. 400 500 600  
700 29 Nanometers 1.

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