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Answers quantitative

analysis where an

unknown sample is

dissolved in an

appropriate solvent,

and the analyte is

converted to an

insoluble form of a

known compound. The

precipitate is then

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collected, dried and weighed.

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Pre-lab Assignment •

Read the lab

thoroughly. • Answer

the pre-lab questions

that appear at the end

of this lab exercise.

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Purpose The purpose
this experiment is to
perform two gas
forming reactions and
determine the actual
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Gravimetric analysis is
a quantitative method
for accurately
determining the

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amount of a substance by selective precipitation of the substance from an aqueous solution. The precipitate is separated from the remaining aqueous solution by filtration and

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Analysis 1. Write the formula and determine the molar mass for the metal carbonate of each of the first three elements in Group 1. 2. If 2.175 g of each metal carbonate were weighed out, how many moles of each substance would you

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Gravimetric analysis is a quantitative method for accurately determining the amount of a substance by selective precipitation of the substance from an aqueous solution. The precipitate is separated from the remaining aqueous solution by filtration and is then weighed. 7:

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Experiment 10

Stoichiometry-
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Prelab Assignment for
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Suppose that 0.323 g
of an unknown sulfate
salt is dissolved in 50
mL of water. The
solution is acidified
with 6M HCl, heated,
and an excess of
aqueous

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quantitative analysis where an unknown sample is dissolved in an appropriate solvent, and the analyte is converted to an insoluble form of a known compound. The precipitate is then collected, dried and weighed.

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Answers analysis is a

quantitative method

for accurately

determining the

amount of a substance

by selective

precipitation of the

substance from an

aqueous solution. The

precipitate is separated

from the remaining

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Questions Gravimetric
Analysis 1. Write the
formula and determine
the molar mass for the
metal carbonate of
each of the first three
elements in Group 1. 2.
If 2.175 Gravimetric

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Answers K_2CO_3 is the limiting reactant 3.

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Experiment 10

Stoichiometry-

Gravimetric Analysis

10- 1 Experiment 10

Stoichiometry-

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Pre-lab Assignment •

Read the lab

thoroughly. • Answer

the pre-lab questions

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that appear at the end of this lab exercise.

Purpose The purpose of this experiment is to perform two gas forming reactions and determine the actual

Experiment 10 Stoichiometry- Gravimetric Analysis

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K_2CO_3 is the limiting reactant 3. Using your answer for Question 2, determine the mass of potassium carbonate needed to fully precipitate all the calcium from a 25 mL sample of 15% calcium chloride.

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Gravimetric Analysis 1. is the most common method for isolating and analyte for gravimetric analysis. 2. Nucleations sites are important to insure the formation of large This is achieved more efficiently by mixing slowly and keeping the solution 3. Other salts with low solubility may generate errors in the analysis.

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